

11-3 Practice Problems

1. Identify the limiting reactant when 1.22-g of O_2 reacts with 1.65-g of H_2 to produce water.

2. Identify the limiting reactant when 4.65-g of Fe reacts with 2.88-g of S to produce FeS .

3. Identify the limiting reactant when 2.87-g of $Mg(OH)_2$ reacts with 12.84-g of HCl to form $MgCl_2$ and water.

4. Identify the limiting reactant when 6.25-g of $AgNO_3$ reacts with 4.12-g of $NaCl$ to form $NaNO_3$ and $AgCl$.

5. Identify the limiting reactant when 7.81-g of HCl reacts with 5.24-g of $NaOH$ to produce $NaCl$ and H_2O .

6. Identify the limiting reactant when 6.33-g of H_2SO_4 reacts with 5.92-g of $NaOH$ to produce Na_2SO_4 and water.

7. Identify the limiting reactant when 43.25-g of CaC_2 reacts with 33.71-g of water to produce $Ca(OH)_2$ and C_2H_2 .

8. Identify the limiting reactant when 63.14-g of $CaCl_2$ reacts with 74.68-g of Na_2CO_3 to produce $CaCO_3$ and $NaCl$.

9. Identify the limiting reactant when 4.687-g of SF_4 reacts with 6.281-g of LiO_2 to produce LiF and SO_2 .

10. If 4.1-g of Cr is heated with 9.3-g of Cl_2 , what mass $CrCl_3$ will be produced?

11. What mass of SO_2 is produced from the reaction between 31.5-g of S_8 and 8.63-g of O_2 ?

12. What mass of SO_2 is produced from the reaction of 12.4-g of SO_2 and 3.45-g of O_2 ?

13. What mass of H_2SO_4 is produced from the reaction of 6.58-g of SO_3 and 1.64-g of H_2O ?

14. What mass of CdS is produced if 8.47-g of cadmium reacts with 2.51-g of Sulfur?

Prentice Hall Chemistry Chapter 11 Practice Problems

**Harold Eugene LeMay, Herbert
Beall, Karen M. Robblee, Douglas C.
Brower**

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Addison-Wesley Chemistry Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, Prentice-Hall Staff, 2001-02 To purchase or download a workbook click on the Purchase or Download button to the left To purchase a workbook enter the desired quantity and click Add to Cart To download a free workbook right click the FREE Download PDF link and save to your computer This will result in a faster download as opposed to left clicking and opening the link

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Lawrence Brown, H. Eugene LeMay, Bruce E. Bursten, Patrick Woodward, Catherine Murphy, 2017-01-03 NOTE This edition features the same content as the traditional text in a convenient three hole punched loose leaf version Books a la Carte also offer a great value this format costs significantly less than a new textbook Before purchasing check with your instructor or review your course syllabus to ensure that you select the correct ISBN Several versions of MyLab tm and Mastering tm platforms exist for each title including customized versions for individual schools and registrations are not transferable In addition you may need a Course ID provided by your instructor to register for and use MyLab and Mastering products For courses in two semester general chemistry Accurate data driven authorship with expanded interactivity leads to greater student engagement Unrivaled problem sets notable scientific accuracy and currency and remarkable clarity have made Chemistry The Central Science the leading general chemistry text for more than a decade Trusted innovative and calibrated the text increases conceptual understanding and leads to greater student success in general chemistry by building on the

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