

# Matter—Properties and Changes

## Section 3.1 Properties of Matter

In your textbook, read about physical properties and chemical properties of matter.

Use each of the terms below just once to complete the passage.

chemical  
density

mass  
properties

physical  
substance

Matter is anything with **(1)** \_\_\_\_\_ and volume. A **(2)** \_\_\_\_\_ is a form of matter with a uniform and unchanging composition. Substances have specific, unchanging **(3)** \_\_\_\_\_ that can be observed. Substances have both physical and chemical properties. **(4)** \_\_\_\_\_ properties can be observed without changing a substance's chemical composition. Color, hardness, and **(5)** \_\_\_\_\_ are examples. Other properties cannot be observed without changing the composition of a substance. These are called **(6)** \_\_\_\_\_ properties. An example is the tendency of iron to form rust when exposed to air.

Label each property as either *physical* or *chemical*.

- \_\_\_\_\_ 7. Chemical formula  $H_2O$
- \_\_\_\_\_ 8. Forms green carbonate when exposed to moist air
- \_\_\_\_\_ 9. Remains unchanged when in the presence of nitrogen
- \_\_\_\_\_ 10. Colorless
- \_\_\_\_\_ 11. Solid at normal temperatures and pressures
- \_\_\_\_\_ 12. Ability to combine with another substance
- \_\_\_\_\_ 13. Melting point
- \_\_\_\_\_ 14. Liquid at normal temperatures and pressures
- \_\_\_\_\_ 15. Boiling point is  $100^{\circ}C$
- \_\_\_\_\_ 16. Conducts electricity
- \_\_\_\_\_ 17. Density is  $\frac{1g}{cm^3}$

# Chemistry Chapter 1 Study Guide

**Victor M. Corman**



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