



**Across:** 3. A system, which can exchange energy, but not matter, with its surroundings. 6. Enthalpy is usually measured on the \_\_\_\_-joules scale. 7. Heat and work are not \_\_\_\_\_ functions. 11. The sample for which energy is studied. 15. The energy of interaction between particles, due to the position of a body (possessing energy) or to the configuration of its parts. 16. A type of change that tends to occur without needing to be driven by an external influence. 17. A process where a substance is converted from its standard state into free atoms. 18. The system and its surroundings. 21. A process, the heat of which is equal in value and opposite in sign to the heat of formation of a given substance. 22. For molar heat capacities of gases,  $C_p = C_v + \underline{\hspace{1cm}}$  (two words). 23. Enthalpy change ( $\Delta H$ ) is the heat absorbed/released for a process at constant \_\_\_\_\_. 24. In the formula,  $\Delta \underline{\hspace{1cm}} = \Delta U + \Delta n \times R \times T$ ,  $\Delta n$  represents the change in the amount of substance of gas in a chemical reaction. 25. Enthalpy of solvation is \_\_\_\_\_ positive. 26. A reaction enthalpy is the sum of the enthalpies of any sequence of reactions (at the same  $T$  and  $P$ ) into which the overall reaction may be divided (two words).

**Down:** 1. Free energy \_\_\_\_\_ with temperature. 2. Standard entropies of the elements are not \_\_\_\_\_. 4. Molar heat capacity is the amount of heat required to raise the temperature of \_\_\_\_\_ of a substance by 1 K (or  $1^\circ\text{C}$ ) (two words). 5. From a more condensed to a less condensed phase,  $\Delta H$  is always \_\_\_\_\_. 8. The standard enthalpy of formation is the heat change when 1 mole of a substance is formed from its \_\_\_\_\_ in their most stable form. 9. The standard enthalpy of combustion is a \_\_\_\_\_ property. 10. In the formula,  $\Delta H = \Delta X + \Delta n \times R \times T$ ,  $X$  represents \_\_\_\_\_ (two words). 12. \_\_\_\_\_ bond enthalpy  $B(A-B)$  is the average energy required to break a given bond type (averaged over a series of related bonds/compounds). 13. A process where a substance is completely burnt in oxygen. 14. Everything outside the system. 19. A process is spontaneous, if free energy is \_\_\_\_\_. 20. The change in internal energy of a system equals the heat released/absorbed by the system and the work done on/by the system (two words).

# **Mark Twain Media Answer Key On Economics**

**Karl Biedenweg, Ph.D.**



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