

# Chemistry

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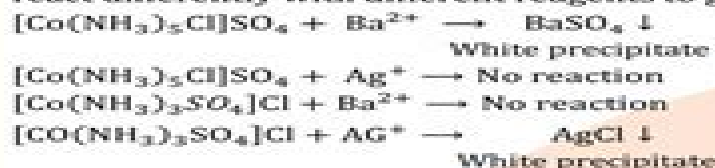
(Chapter 9) (Coordination Compound) (Intext Questions)  
(Class 12)

## Question 9.4:

Give evidence that  $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{SO}_4$  and  $[\text{Co}(\text{NH}_3)_5\text{SO}_4]\text{Cl}$  are ionization isomers.

### Answer 9.4:

When ionization isomers are dissolved in water, they ionize to give different ions. The ions then react differently with different reagents to give different products.



## Question 9.5:

Explain on the basis of valence bond theory that  $[\text{Ni}(\text{CN})_4]^{2-}$  ion with square planar structure is diamagnetic and the  $[\text{NiCl}_4]^{2-}$  ion with tetrahedral geometry is paramagnetic.

### Answer 9.5:

Ni is in the +2 oxidation state i.e., in  $d^8$  configuration.



There are 4  $\text{CN}^-$  ions. Thus, it can either have a tetrahedral geometry or square planar geometry. Since  $\text{CN}^-$  ion is a strong field ligand, it causes the pairing of unpaired 3d electrons.



It now undergoes  $dsp^2$  hybridization. Since all electrons are paired, it is diamagnetic. In case of  $[\text{NiCl}_4]^{2-}$ ,  $\text{Cl}^-$  ion is a weak field ligand. Therefore, it does not lead to the pairing of unpaired 3d electrons. Therefore, it undergoes  $sp^3$  hybridization.



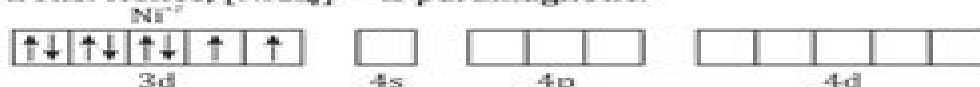
Since there are 2 unpaired electrons in this case, it is paramagnetic in nature.

## Question 9.6:

$[\text{NiCl}_4]^{2-}$  is paramagnetic while  $[\text{Ni}(\text{CO})_4]$  is diamagnetic though both are tetrahedral. Why?

### Answer 9.6:

Though both  $[\text{NiCl}_4]^{2-}$  and  $[\text{Ni}(\text{CO})_4]$  are tetrahedral, their magnetic characters are different. This is due to a difference in the nature of ligands.  $\text{Cl}^-$  is a weak field ligand and it does not cause the pairing of unpaired 3d electrons. Hence,  $[\text{NiCl}_4]^{2-}$  is paramagnetic.



In  $[\text{Ni}(\text{CO})_4]$ , Ni is in the zero oxidation state i.e., it has a configuration of  $3d^8 4s^2$ .

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