

Detection of Ions in Solutions Using Acid/Base Chemistry: A Quality Control Test

Objective: This lab focuses on the detection of ions using titration as an analysis tool. You will standardize NaOH and HCl solutions so that you know the exact concentration and then prepare samples of common household items in order to determine the amount of calcium in Tang[®], Mg(OH)₂ in Milk of Magnesia, etc. You will learn to prepare samples of a specified concentration, learn about acids and bases through the use of titrations and learn how to detect endpoints using different indicators. You will become adept at measuring pH with both a pH meter and indicator paper. You will then conduct quality control testers and determine if the label on a bottle of over-the-counter product actually contains the percentage of compound that it advertises.

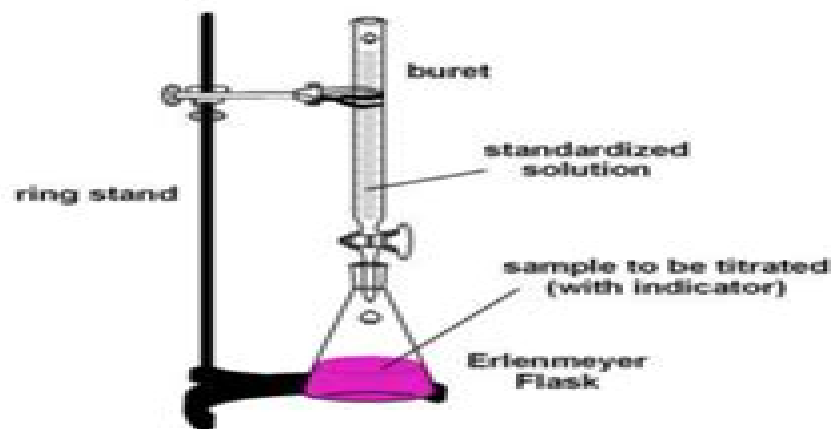
What is a Titration?

A titration is an analytical procedure used to determine the concentration of a sample by reacting it with a standard solution. One type of titration uses a neutralization reaction, in which an acid and a base react to produce a salt and water.



In equation 1, the acid is HCl (called hydrochloric acid) and the base is NaOH (called sodium hydroxide). When the acid and base react, they form NaCl (sodium chloride), which is also known as table salt. The titration proceeds until the equivalence point is reached, where the number of moles of acid is equal to the number of moles of base. This point is usually marked by observing a color change in an added indicator.

In a titration, the standard solution goes in a buret, which is a piece of glassware used to measure the volume of solvent to approximately 0.1 mL of accuracy. The solution that you are titrating goes in an Erlenmeyer flask, which should be large enough to accommodate both your sample and the standard solution you are adding.



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Quantitative Chemical Analysis Daniel C. Harris, Chuck Lucy, 2015-05-29 The gold standard in analytical chemistry Dan Harris Quantitative Chemical Analysis provides a sound physical understanding of the principles of analytical chemistry and their applications in the disciplines

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OECD Guidelines for the Testing of Chemicals, Section 1 Test No. 122: Determination of pH, Acidity and Alkalinity OECD, 2013-07-26 This Test Guideline describes the procedure for the electronic determination of pH of an undiluted aqueous solution or dispersion the pH of a dilution of a solution or dispersion in water or the pH of a chemical diluted to end use concentration

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Fundamentals of Analytical Chemistry Douglas A. Skoog, Donald M. West, Stanley R. Crouch, F. James Holler, 2013-01-01 Known for its readability and systematic rigorous approach this fully updated FUNDAMENTALS OF ANALYTICAL CHEMISTRY 9E International Edition offers extensive coverage of the principles and practices of analytic chemistry and consistently shows students its applied nature The book s award winning authors begin each chapter with a story and photo of how analytic chemistry is applied in industry medicine and all the sciences To further reinforce student learning a wealth of dynamic photographs by renowned chemistry photographer Charlie Winters appear as chapter openers and throughout the text Incorporating Excel spreadsheets as a problem solving tool the Ninth Edition is enhanced by a chapter on Using Spreadsheets in Analytical Chemistry updated spreadsheet summaries and problems an Excel Shortcut Keystrokes for the PC insert card and a supplement by the text authors EXCEL APPLICATIONS FOR ANALYTICAL CHEMISTRY which integrates this important aspect of the study of analytical chemistry into the book s already rich pedagogy New to this edition is OWL an online homework and assessment tool that includes the Cengage YouBook a fully

customizable and interactive eBook which enhances conceptual understanding through hands on integrated multimedia interactivity

Guidance for Preparing Standard Operating Procedures (SOPs).,2001 **Laboratory Experiments for Brown and LeMay, Chemistry, the Central Science** John Henry Nelson,1985 **Holt Chemfile B Microscale Exp/Te 2006** Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2006 **Aqueous Acid-base Equilibria and Titrations** Robert De Levie,1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an out moded way This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved

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LabWorks Joseph R. Crook,1997 Quantifying Uncertainty in Analytical Measurement Eurachem/CITAC Working Group,S. L. R. Ellison,M. Rosslein,Alexander Williams,2000-01-01 Standard Methods for the Examination of Water and Wastewater American Public Health Association,1971 Martin's Physical Pharmacy and Pharmaceutical Sciences Alfred N. Martin,Patrick J. Sinko,2011 Martin s Physical Pharmacy and Pharmaceutical Sciences is considered the most comprehensive text available on the application of the physical chemical and biological principles in the pharmaceutical sciences It helps students teachers researchers and industrial pharmaceutical scientists use elements of biology physics and chemistry in their work and study Since the first edition was published in 1960 the text has been and continues to be a required text for the core courses of Pharmaceutics Drug Delivery and Physical Pharmacy The Sixth Edition features expanded content on drug delivery solid oral dosage forms pharmaceutical polymers and pharmaceutical

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