

10**CHEMICAL QUANTITIES****Practice Problems**

On a piece of paper, solve the following problems.

YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

1. What is the molar mass of sucrose ($C_{12}H_{22}O_{11}$)?
2. What is the molar mass of each of the following compounds?
 - a. phosphorus pentachloride (PCl_5)
 - b. uranium hexafluoride (UF_6)
3. Calculate the molar mass of each of the following ionic compounds:
 - a. $KMnO_4$
 - b. $Ca_3(PO_4)_2$
4. How many moles is 3.52×10^{24} molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen (O_2)?

SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. What is the molar mass of each of the following compounds?
 - a. $C_6H_{12}O_6$
 - b. $NaHCO_3$
 - c. C_2H_2
 - d. KNH_4SO_4
2. Calculate the mass in grams of each of the following:
 - a. 8.0 mol lead oxide (PbO)
 - b. 0.75 mol hydrogen sulfide (H_2S)
 - c. 0.00100 mol silicon tetrahydride (SiH_4)
 - d. 1.50×10^{-2} mol molecular oxygen (O_2)
 - e. 2.30 mol ethylene glycol ($C_2H_6O_2$)
3. How many grams are in 1.73 mol of dinitrogen pentoxide (N_2O_5)?
4. How many grams are in 0.658 mol of calcium phosphate [$Ca_3(PO_4)_2$]?
5. Calculate the number of moles in each of the following:
 - a. 0.50 g sodium bromide ($NaBr$)
 - b. 13.5 g magnesium nitrate [$Mg(NO_3)_2$]
 - c. 1.02 g magnesium chloride ($MgCl_2$)
 - d. 0.00100 g monochloromethane (CH_3Cl)
 - e. 1.50×10^{-3} g propylene glycol [$C_3H_8(OH)_2$]
6. A chemist plans to use 435.0 grams of ammonium nitrate (NH_4NO_3) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ($C_{20}H_{24}N_2O_2$). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane (CH_4) gas?
9. How many moles is 135 L of ammonia (NH_3) gas at STP?

Chemistry Chapter 16 Solutions Workbook Answers

Arshad Iqbal



Chemistry Chapter 16 Solutions Workbook Answers

Organic Chemistry K. Peter C. Vollhardt, Neil Eric Schore, 2007 This textbook provides students with a framework for organizing their approach to the course dispelling the notion that organic chemistry is an overwhelming shapeless body of facts

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Solutions Manual with Answers to All Questions, Analytical Chemistry, Principles and Techniques Larry G. Hargis, 1987-12

Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools

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Basic Concepts of Chemistry, Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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needed to actually solve problems **Chemistry, Student Solutions Manual** James N. Spencer, Alexander Grushow, George M. Bodner, Lyman H. Rickard, 2008-01-09 This innovative text provides a 15 chapter introduction to the fundamental concepts of chemistry The material is then supplemented by special topics at the end of each chapter

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 - Chapter 1: Introduction to Chemistry Chapter 16 Solutions Workbook Answers
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Table of Contents Chemistry Chapter 16 Solutions Workbook Answers

1. Understanding the eBook Chemistry Chapter 16 Solutions Workbook Answers
 - The Rise of Digital Reading Chemistry Chapter 16 Solutions Workbook Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 16 Solutions Workbook Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 16 Solutions Workbook Answers
 - Personalized Recommendations
 - Chemistry Chapter 16 Solutions Workbook Answers User Reviews and Ratings
 - Chemistry Chapter 16 Solutions Workbook Answers and Bestseller Lists
5. Accessing Chemistry Chapter 16 Solutions Workbook Answers Free and Paid eBooks
 - Chemistry Chapter 16 Solutions Workbook Answers Public Domain eBooks
 - Chemistry Chapter 16 Solutions Workbook Answers eBook Subscription Services
 - Chemistry Chapter 16 Solutions Workbook Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 16 Solutions Workbook Answers eBook Formats
 - ePub, PDF, MOBI, and More

- Chemistry Chapter 16 Solutions Workbook Answers Compatibility with Devices
 - Chemistry Chapter 16 Solutions Workbook Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 16 Solutions Workbook Answers
 - Highlighting and Note-Taking Chemistry Chapter 16 Solutions Workbook Answers
 - Interactive Elements Chemistry Chapter 16 Solutions Workbook Answers
 8. Staying Engaged with Chemistry Chapter 16 Solutions Workbook Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 16 Solutions Workbook Answers
 9. Balancing eBooks and Physical Books Chemistry Chapter 16 Solutions Workbook Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 16 Solutions Workbook Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Chapter 16 Solutions Workbook Answers
 - Setting Reading Goals Chemistry Chapter 16 Solutions Workbook Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Chapter 16 Solutions Workbook Answers
 - Fact-Checking eBook Content of Chemistry Chapter 16 Solutions Workbook Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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