

HALF YEARLY EXAMINATION — 2014-15

CHEMISTRY

Class – XI

Time Allowed : 3 Hrs.

Maximum Marks : 70

General Instructions :

- (i) Attempt all questions.
- (ii) Q 1-5 are very short answer type questions carrying 1 mark each.
- (iii) Q 6-10 are short answer type questions carrying 2 marks each.
- (iv) Q 11-22 are also short answer type questions carrying 3 marks each.
- (v) Q 23 is a value based question carrying 4 marks.
- (vi) Q 24-26 are long answer type questions carrying 5 marks each.
- (vii) There is no overall choice in the paper. However, there is an internal choice in some questions.
- (viii) Use of calculator is prohibited. Log tables can be used if necessary.

1. Give an example of a molecule which has intramolecular hydrogen bonding.
2. Write the electronic configuration of La ($z = 57$).
3. Which of the following species has the largest and smallest size? Mg, Mg^{2+} , Al, Al^{3+}
4. What is the significance of van der Waals parameters 'a' and 'b'?
5. What will happen to the wavelength associated with the moving particle if its velocity is doubled?
6. Two bulbs B_1 and B_2 of equal capacity contains 10g oxygen (O_2) and ozone (O_3) respectively. Which bulb will have greater number of oxygen atoms?
(Atomic mass of oxygen = 16 u and Avogadro's number = 6.022×10^{23})
7. Using ideal gas equation, show that at a given temperature, density of a gas is directly proportional to the pressure of the gas.
8. Derive a relation between C_p and C_v for one mole of an ideal gas.
9. Explain dispersion forces with the help of a suitable example.

OR

The Maxwell-Boltzmann distribution curve for a gas at three different temperatures is shown below. Answer the questions based on it.

- (a) The value of the most probable speed (c^*) obtained at temperature T_3 is greater than that obtained at temperature T_1 and T_2 . Justify.

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S Ashworth



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