

therefore +3 each.

nitrogen(III) oxide

- Carbon is +4 and each oxygen is -2.
  - Carbon is -4 and each hydrogen is +1.
  - Each carbon is 0, each hydrogen is +1, and each oxygen is -2.
  - Each carbon is -8/3 and each hydrogen is +1.
- Each iron is +3 and each oxygen is -2.
  - Nitrogen is +4 and each oxygen is -2.
  - Chromium is +4 and each oxygen is -2.

## SECTION 3

### SHORT ANSWER

- True
  - False
  - True
- 10 mol of calcium, 20 mol of nitrogen, 60 mol of oxygen
- 0.250 mol
  - $1.50 \times 10^{23}$  molecules
  - $4.51 \times 10^{23}$  carbon atoms
  - 10.1%
- 52.9%
  - 2100 lb
- 20 g
  - 0.17 mol
  - 4.0 mol

## Section 4

### SHORT ANSWER

- $\text{CH}_3\text{O}_2$
  - $\text{N}_2\text{O}_5$
  - $\text{HgCl}$
  - $\text{CH}_2$
- $\text{C}_4\text{H}_8$
- $\text{Na}_2\text{S}_2\text{O}_3$
  - neither
- 36%
  - 5
  - The second heating is to ensure that all the water in the sample has been driven off. If the mass is less after the second heating, water was still present after the first heating.
- $\text{CF}_2$
  - $\text{C}_4\text{F}_8$
- $\text{CuClO}_3$
  - copper(I) chlorate

## MIXED REVIEW

### SHORT ANSWER

- $\text{CuCO}_3$
  - $\text{Na}_2\text{SO}_5$
  - $(\text{NH}_4)_3\text{PO}_4$
  - $\text{SnS}_2$
  - $\text{HNO}_2$
- magnesium perchlorate
  - iron(II) nitrate
  - iron(III) nitrite
  - cobalt(II) oxide
  - nitrogen(V) oxide

## Chapter 8

### SECTION 1

#### SHORT ANSWER

- d
  - a
  - b
  - f
  - e
  - c
- 8, 4, 9
- 12 atoms
  - 16 atoms
  - 51 atoms
  - $3 \times 10^{24}$  atoms
- $2\text{Al}(s) + 3\text{CuF}_2(aq) \rightarrow 2\text{AlF}_3(aq) + 3\text{Cu}(s)$
- $\text{NaCl}(aq) + \text{AgNO}_3(aq) \rightarrow \text{AgCl}(s) + \text{NaNO}_3(aq)$
- $2\text{NaHCO}_3(s) \xrightarrow{\Delta} \text{Na}_2\text{CO}_3(s) + \text{H}_2\text{O}(g) + \text{CO}_2(g)$
  - When solid sodium hydrogen carbonate (bicarbonate) is heated, it decomposes into solid sodium carbonate while releasing carbon dioxide gas and water vapor.
- It is balanced but incorrect. In two of the formulas the subscripts were changed, which changed the compounds involved. Water is not  $\text{H}_2\text{O}$ , and sodium hydroxide is not  $\text{Na}_2\text{OH}$ . The correct balanced equation is  $2\text{NaOH} + \text{H}_2\text{S} \rightarrow \text{Na}_2\text{S} + 2\text{H}_2\text{O}$ .
- 30 mol
  - 40 mol

## Section 2

### SHORT ANSWER

- c
  - d

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**Richard Bailey**



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