

## Worksheet: Writing and Balancing Chemical Reactions

1. Balance the following equations and indicate the type of reaction as formation, decomposition, single replacement, double replacement, hydrocarbon combustion, or other.

- $\text{Cu}_{(s)} + \text{O}_{2(g)} \rightarrow \text{CuO}_{(s)}$
- $\text{H}_2\text{O}_{(l)} \rightarrow \text{H}_2(g) + \text{O}_2(g)$
- $\text{Fe}_{(s)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2(g) + \text{Fe}_3\text{O}_4(s)$
- $\text{AsCl}_3(aq) + \text{H}_2\text{S}(aq) \rightarrow \text{As}_2\text{S}_3(s) + \text{HCl}(aq)$
- $\text{CuSO}_4 + 5 \text{H}_2\text{O}_{(l)} \rightarrow \text{CuSO}_4 \cdot 5\text{H}_2\text{O}_{(s)}$
- $\text{Fe}_2\text{O}_3(s) + \text{H}_2(g) \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- $\text{CaCO}_3(s) \rightarrow \text{CaO}_{(s)} + \text{CO}_2(g)$
- $\text{Fe}_{(s)} + \text{S}_8(s) \rightarrow \text{FeS}_{(s)}$
- $\text{H}_2\text{S}(aq) + \text{KOH}(aq) \rightarrow \text{H}_2\text{O}_{(l)} + \text{K}_2\text{S}(aq)$
- $\text{NaCl}_{(s)} \rightarrow \text{Na}_{(s)} + \text{Cl}_2(g)$
- $\text{Al}_{(s)} + \text{H}_2\text{SO}_4(aq) \rightarrow \text{H}_2(g) + \text{Al}_2(\text{SO}_4)_3(aq)$
- $\text{H}_3\text{PO}_4(aq) + \text{NH}_4\text{OH}(aq) \rightarrow \text{H}_2\text{O}_{(l)} + (\text{NH}_4)_3\text{PO}_4(aq)$
- $\text{C}_3\text{H}_8(g) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}_{(l)}$
- $\text{Al}_{(s)} + \text{O}_2(g) \rightarrow \text{Al}_2\text{O}_3(s)$
- $\text{CH}_4(g) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}_{(l)}$
- $\text{K}_2\text{SO}_4(aq) + \text{BaCl}_2(aq) \rightarrow \text{KCl}(aq) + \text{BaSO}_4(s)$
- $\text{C}_5\text{H}_{12}(l) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}_{(l)}$
- $\text{Ca}(\text{OH})_2(aq) + \text{NH}_4\text{Cl}(aq) \rightarrow \text{NH}_4\text{OH}(aq) + \text{CaCl}_2(aq)$
- $\text{V}_2\text{O}_5(s) + \text{Ca}_{(s)} \rightarrow \text{CaO}_{(s)} + \text{V}_{(s)}$
- $\text{Na}_{(s)} + \text{ZnI}_2(aq) \rightarrow \text{NaI}(aq) + \text{Zn}_{(s)}$
- $\text{C}_7\text{H}_6\text{O}_3(l) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}_{(l)}$
- $\text{Ca}_{(s)} + \text{N}_2(g) \rightarrow \text{Ca}_3\text{N}_2(s)$
- $\text{Fe}_2\text{O}_3(s) + \text{H}_2(g) \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- $\text{C}_{15}\text{H}_{30}(l) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}_{(l)}$
- $\text{BN}_{(s)} + \text{F}_2(g) \rightarrow \text{BF}_3(g) + \text{N}_2(g)$
- $\text{C}_{12}\text{H}_{26}(l) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}_{(l)}$

2. Predict the product(s) along with the states, indicate the type of reaction, and balance the following chemical reactions.

- A solution of lead (II) nitrate is mixed with a solution of sodium iodide.
- Solid zinc sulfide reacts with oxygen in the air.
- Liquid butane ( $\text{C}_4\text{H}_{10}(l)$ ) is used as a fuel to ignite a lighter.
- Barium hydroxide solution is neutralized by adding hydrochloric acid ( $\text{HCl}(aq)$ ).
- Copper metal is placed in a solution of silver nitrate.
- Sulfur burns in oxygen to make sulfur dioxide gas.
- A solution of aluminum sulfate is mixed with a solution of calcium hydroxide.
- Zinc metal is placed in sulfuric acid ( $\text{H}_2\text{SO}_4(aq)$ ).
- Aluminum powder is placed in a container filled with chlorine gas.
- Sucrose undergoes cellular respiration.

# Balancing Equations With Formulas Names Ws Answers

**Pamphlet Master**



# Balancing Equations With Formulas Names Ws Answers

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are described in the preface to help instructors transition to the second edition **Te HS&T J** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004-02 *CK-12 Chemistry - Second Edition* CK-12 Foundation, 2011-10-14 CK 12 Foundation's Chemistry Second Edition FlexBook covers the following chapters Introduction to Chemistry scientific method history Measurement in Chemistry measurements formulas Matter and Energy matter energy The Atomic Theory atom models atomic structure sub atomic particles The Bohr Model of the Atom electromagnetic radiation atomic spectra The Quantum Mechanical Model of the Atom energy standing waves Heisenberg Schrodinger The Electron Configuration of Atoms Aufbau principle electron configurations Electron Configuration and the Periodic Table electron configuration position on periodic table Chemical Periodicity atomic size ionization energy electron affinity Ionic Bonds and Formulas ionization ionic bonding ionic compounds Covalent Bonds and Formulas nomenclature electronic molecular geometries octet rule polar molecules The Mole Concept formula stoichiometry Chemical Reactions balancing equations reaction types Stoichiometry limiting reactant equations yields heat of reaction The Behavior of Gases molecular structure properties combined gas law universal gas law Condensed Phases Solids and Liquids intermolecular forces of attraction phase change phase diagrams Solutions and Their Behavior concentration solubility colligate properties dissociation ions in solution Chemical Kinetics reaction rates factors that affect rates Chemical Equilibrium forward reverse reaction rates equilibrium constant Le Chatelier's principle solubility product constant Acids Bases strong weak acids and bases hydrolysis of salts pH Neutralization dissociation of water acid base indicators acid base titration buffers Thermochemistry bond breaking formation heat of reaction formation Hess law entropy Gibb's free energy Electrochemistry oxidation reduction electrochemical cells Nuclear Chemistry radioactivity nuclear equations nuclear energy Organic Chemistry straight chain aromatic hydrocarbons functional groups Chemistry Glossary *Balanced Scorecards and Operational Dashboards with Microsoft Excel* Ron Person, 2010-08-26 Created in Excel balanced scorecards enable you to monitor operations and tactics while operational dashboards is a set of indicators regarding the state of a business metric or process both features are in high demand for many large organizations This book serves as the first guide to focus on combining the benefits of balanced scorecards operational dashboards performance managements and data visualization and then implement them in Microsoft Excel

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