

59. Alpha particles with their high ionizing charge but low penetrating ability can best be shielded by?

a. paper

b. clothing

c. water

d. lead

Alpha α particles - paper is a sufficient shield

Beta β particles - heavy clothing is sufficient to shield

Gamma rays - lead is required to shield

↔ = polar

↕ = polar

↗ = non-polar

↔ = non-polar

↕ = non-polar

60. The half-life of Se-121 is 10 days. If you started with 40 grams of this isotope, how much would you have left after 30 days?

a. 10 g

b. none

c. 15 g

d. 5 g

Mass of Se-121	Number of half-lives	Time
40 g	0	0
20 g	1	10 days
10 g	2	20 days
5 g	3	30 days

⬆ = particles go faster

⬇ = particles go slower

PERIODIC TRENDS

ATOMIC RADIUS: decreases →
increases ↓

METALLIC CHARACTER:

decreases ↓
increases →

IONIZATION ⬆ decreases ↓
increases →

A → mass # (# protons + # neutrons)
Z → atomic # (# of protons)
X → chemical symbol

• Isotopes = atoms with same # of protons, but different # of neutrons.
EXAMPLE: Ne-20 → $^{20}_{10}\text{Ne}$
Ne-21 → $^{21}_{10}\text{Ne}$
Ne-22 → $^{22}_{10}\text{Ne}$

• Dilution Equation $M_1 V_1 = M_2 V_2$

• $\text{mass\%} = \frac{\text{mass solute}}{\text{mass of solution (solute + solvent)}} \times 100\%$

• $\text{Molarity (M)} = \frac{\text{mol solute}}{\text{L solution}}$

• $\text{molality (m)} = \frac{\text{mol solute}}{\text{kg solvent}}$

• #protons - #electrons = Charge #

• #neutrons = A (mass #) - Z (atomic #)

• if no charge, the #protons + #electrons = same

• if + (positive) charge (cation) = #electrons is less than #protons

• if - (negative) charge (anion) = #electrons is greater than #protons

• Protons ≠ Charge → they give identity to the element.

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Table of Contents Acs High School Chemistry Exam Study Guide

1. Understanding the eBook Acs High School Chemistry Exam Study Guide
 - The Rise of Digital Reading Acs High School Chemistry Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Acs High School Chemistry Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acs High School Chemistry Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acs High School Chemistry Exam Study Guide
 - Personalized Recommendations
 - Acs High School Chemistry Exam Study Guide User Reviews and Ratings

- Acs High School Chemistry Exam Study Guide and Bestseller Lists
- 5. Accessing Acs High School Chemistry Exam Study Guide Free and Paid eBooks
 - Acs High School Chemistry Exam Study Guide Public Domain eBooks
 - Acs High School Chemistry Exam Study Guide eBook Subscription Services
 - Acs High School Chemistry Exam Study Guide Budget-Friendly Options
- 6. Navigating Acs High School Chemistry Exam Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Acs High School Chemistry Exam Study Guide Compatibility with Devices
 - Acs High School Chemistry Exam Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acs High School Chemistry Exam Study Guide
 - Highlighting and Note-Taking Acs High School Chemistry Exam Study Guide
 - Interactive Elements Acs High School Chemistry Exam Study Guide
- 8. Staying Engaged with Acs High School Chemistry Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acs High School Chemistry Exam Study Guide
- 9. Balancing eBooks and Physical Books Acs High School Chemistry Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acs High School Chemistry Exam Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acs High School Chemistry Exam Study Guide
 - Setting Reading Goals Acs High School Chemistry Exam Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acs High School Chemistry Exam Study Guide
 - Fact-Checking eBook Content of Acs High School Chemistry Exam Study Guide
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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